SCH 3U

**DESIGN OF CHEMICAL REACTION DEMONSTRATION**

Purpose:

To clearly write a procedure that can be successfully followed by another student; to choose reactants that will result in a single displacement and a double displacement reaction; to critique the procedure of another student.

Procedure:

Write the title, purpose, materials and procedure of an experiment that will have two parts: the demonstration of a single displacement reaction, and the demonstration of a double displacement reaction. Choose reactants from the following list without duplicating the other experiments completed in the last few classes. When making your choice, write the expected reaction in your notes and check the activity series and solubility table to ensure that a reaction will occur. Do not include the complete reaction in your procedure.

The procedure must be in the imperative tense, in numbered steps. Use proper terminology and write in logical order.

Available Reactants:

|  |  |  |
| --- | --- | --- |
| Metals | Solutions | |
| magnesium | copper(II) chloride | zinc nitrate |
| nickel | sodium hydroxide | silver nitrate |
| aluminum | iron(III) chloride | potassium chloride |
| lead | lead(II) acetate | copper(II) nitrate |
| zinc | lead(II) nitrate | sodium iodide |
| tin | potassium iodide | hydrochloric acid |
| copper | zinc sulfate | sulfuric acid |