

**IONIC COMPOUNDS:
Names and Formulas**

1. Write the formulas for the following compounds.

- | | |
|------------------------------|-------------------------------|
| a. magnesium oxide _____ | k. tin (II) fluoride _____ |
| b. aluminum nitride _____ | l. lead (IV) nitride _____ |
| c. potassium sulfide _____ | m. iron (III) chloride _____ |
| d. calcium bromide _____ | n. copper (I) oxide _____ |
| e. aluminum sulfide _____ | o. _____ |
| f. beryllium oxide _____ | p. mercury (II) oxide _____ |
| g. strontium phosphide _____ | q. tin (IV) iodide _____ |
| h. sodium fluoride _____ | r. _____ |
| i. lithium selenide _____ | s. cobalt (II) sulphide _____ |
| j. barium oxide _____ | t. tin (IV) sulphide _____ |

2. Write the names for the following compounds.

- | | |
|----------------------------------|--------------------------------|
| a. Li_2O _____ | k. PbS _____ |
| b. AlCl_3 _____ | l. SnO_2 _____ |
| c. MgS _____ | m. NiO _____ |
| d. CaF_2 _____ | n. CuI_2 _____ |
| e. Al_2O_3 _____ | o. PbCl_4 _____ |
| f. BeF_2 _____ | p. FeP _____ |
| g. K_3P _____ | q. AuBr_3 _____ |
| h. Mg_3P_2 _____ | r. Hg_2S _____ |
| i. CaO _____ | s. _____ |
| j. Ag_2S _____ | t. MnO_2 _____ |

IONIC COMPOUNDS: Names and Formulas

1. Write the formulas for the following compounds.

- a. magnesium oxide MgO
- b. aluminum nitride AlN
- c. potassium sulfide K₂S
- d. calcium bromide CaBr₂
- e. aluminum sulfide Al₂S₃
- f. beryllium oxide BeO
- g. strontium phosphide Sr₃P₂
- h. sodium fluoride NaF
- i. lithium selenide Li₂Se
- j. barium oxide BaO

- k. tin (II) fluoride SnF₂
- l. lead (IV) nitride Pb₃N₄
- m. iron (III) chloride FeCl₃
- n. copper (I) oxide Cu₂O
- o. ~~_____~~
- p. mercury (II) oxide HgO
- q. tin (IV) iodide SnI₄
- r. ~~_____~~
- s. cobalt (II) sulphide CoS
- t. tin (IV) sulphide SnS₂

2. Write the names for the following compounds.

- a. Li₂O lithium oxide
- b. AlCl₃ aluminum chloride
- c. MgS magnesium sulfide
- d. CaF₂ calcium fluoride
- e. Al₂O₃ aluminum oxide
- f. BeF₂ beryllium fluoride
- g. K₃P potassium phosphide
- h. Mg₃P₂ magnesium phosphide
- i. CaO calcium oxide
- j. Ag₂S silver sulfide

- k. PbS lead(II) sulfide
- l. SnO₂ tin(IV) oxide
- m. NiO nickel(II) oxide
- n. CuI₂ copper(II) iodide
- o. PbCl₄ lead(IV) chloride
- p. FeP iron(III) phosphide
- q. AuBr₃ gold(III) bromide
- r. Hg₂S mercury(I) sulfide
- s. ~~_____~~
- t. MnO₂ manganese(IV) oxide